

SECTION 1 continued

10. If electrons involved in bonding spend most of the time closer to one atom rather than the other, the bond is _____.

11. If a bond's character is more than 50% ionic, then the bond is called a(n) _____.

12. A bond's character is more than 50% ionic if the electronegativity difference between the two atoms is greater than _____.

13. Write the formula for an example of each of the following compounds:

_____ a. nonpolar covalent compound

_____ b. polar covalent compound

_____ c. ionic compound

14. Describe how a covalent bond holds two atoms together.

15. What property of the two atoms in a covalent bond determines whether or not the bond will be polar?

16. How can electronegativity be used to distinguish between an ionic bond and a covalent bond?

17. Describe the electron distribution in a polar-covalent bond and its effect on the partial charges of the compound.
